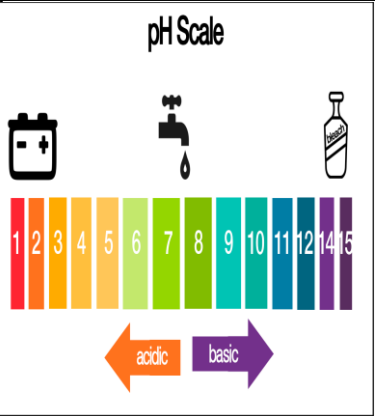

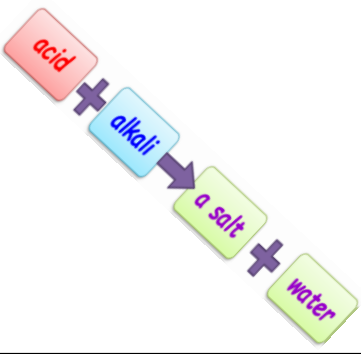
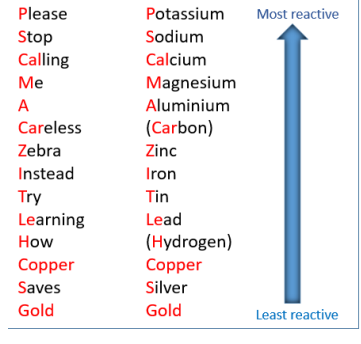


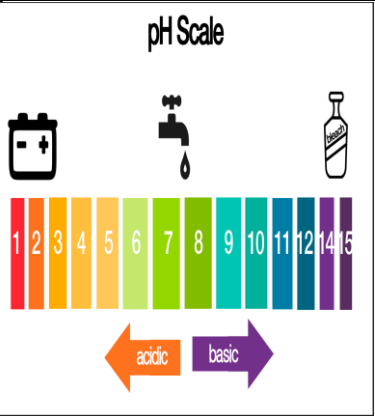

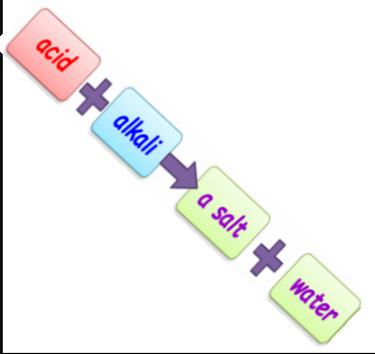
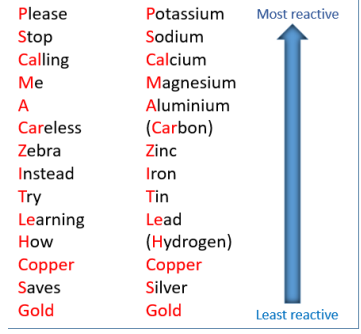
Rationale

To understand how chemicals can be classified based on their properties including acids, alkalis and neutral substances. This is a fundamental topic in Chemistry.

Diagrams	Keywords	Definitions
<p>pH Scale</p> 	<p>Acid</p> <p>Alkali</p> <p>Neutral</p> <p>pH scale</p> <p>Base</p>	<p>Corrosive substance which has a pH lower than 7.</p> <p>A base which is soluble in water.</p> <p>Neither a acid or alkali, has pH of 7.</p> <p>The pH scale shows how acidic a substance is. It can be measured using a pH meter which gives a numerical value.</p> <p>A substance that has pH between 8-14.</p>
	<p>Indicator</p> <p>Universal indicator</p>	<p>A substance which changes colour when present in acid, alkali or neutral substance.</p> <p>A substances that changes colour to represent the pH scale.</p>
	<p>Chemical reaction</p> <p>Neutralisation</p> <p>Word equations</p>	<p>A change that occurs when you chemicals are reacted together, making a new product.</p> <p>When an acid and alkali are mixed together, they undergo a reaction called neutralisation.</p> <p>An equation in which only the names of the reactants and products are used to model a reaction.</p>
	<p>Reactivity series</p> <p>Metal + Acid</p> <p>Metal + Water</p>	<p>The reactivity of a metal. Some metals are very reactive. They easily take part in chemical reactions to make new substances.</p> <p>When a acid is reacted with the metal, the products formed is <u>metal salt</u> + Hydrogen</p> <p>If the metal is a reactive metal, when reacted with water it will form <u>metal oxide</u> + Hydrogen</p>
<p>Hydrochloric Acid = CHLORIDE Salt</p> <p>Nitric Acid = NITRATE Salt</p> <p>Sulfuric Acid = SULFATE Salt</p>	<p>Salts</p>	<p>When a alkali reacts with a acid it form salt. The type of salt depends on the acid.</p>

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